

# Self Assessment Test

## Chemistry

### Std IX

**F.M. 20**

**Time: 45 minutes**

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Q1. Write the balanced chemical equations of the following reactions. 1×5 = 5 marks

(i) Ammonia + Oxygen → Nitric Oxide + Water

(ii) Potassium Bicarbonate + Sulphuric Acid → Potassium Sulphate + Carbon Dioxide + Water

(iii) Ammonium Chloride + Calcium Hydroxide → Calcium Chloride + Ammonia + Water

(iv) Aluminium + Sodium hydroxide + Water → Sodium meta aluminate + Hydrogen

(v) Zinc Sulphide + Oxygen → Zinc Oxide + Sulphur Dioxide

Q2. Calculate the relative molecular masses of : 1×3 = 3 marks

(i)  $(\text{NH}_4)_2\text{SO}_4$

(ii)  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$

(iii)  $\text{CH}_3\text{COONa}$

Given the Mass of : N =14, H= 1, S =32, O = 16, Cu = 63.5, C =12, Na = 23

Q3. Calculate the percentage of Phosphorus in  $\text{Ca}(\text{H}_2\text{PO}_4)_2$  2 marks

Given mass of : Ca = 40, H = 1, P = 31, O = 16

Q4. Write a chemical reaction involving: 1× 3 = 3 marks

(a) Evolution of gas

(b) Change of colour

(c) Formation of precipitate

Q5. What do you observe when: 1×3 = 3 marks

(a) Lead nitrate is heated

(b) Barium chloride is added to Sodium sulphate solution

(c)  $\text{H}_2\text{S}$  gas is passed through Copper sulphate solution

Q6. Define Displacement reaction. Give one example of it. 2 marks

Q7. Write two limitations of a chemical reaction. 2 marks